Safety Data Sheet
Liquid Foot Care

Section 1
Identification of the Substance or Preparation

1.1Product Name:- Liquid Foot Care
Other Names:- Milk of Lime.

1.2Address/Tel:
Bolshaws Industrial Powders, Harrop House, Rainow, Macclesfield SK10 5UU

1.3 In an Emergency:- Dial 999
For specialist advice (transport emergency)
Tel 0044(0)1625 572416

Section 2
Composition/Information on Ingredients

Suspension of Calcium Hydroxide Ca $(OH)_2$ in water. Small quantities of calcium carbonate, silica and oxides of magnesium, aluminium and iron and other trace elements.

Hazardous Ingredient – Calcium Hydroxide
R38, R41 See detail under section 15
CAS No:- 1305-62-0
EINECS No:- 215-137-3

Section 3
Hazard Identification

The product can produce severe skin damage in humans, (alkaline burns), especially if prolonged skin contact takes place.

Section 4
First Aid Measures

4.1Emergency first aid procedures
Skin – Carefully and gently brush the contaminated body surfaces in order to remove all traces of product. Wash affected area immediately with plenty of water. Remove contaminated clothing. If necessary seek medical advice.

Eyes – Irrigate eyes immediately with plenty of water and seek medical advice.

Inhalation – Not applicable

Ingestion – Wash mouth with water and drink copious quantities of water. Do not induce vomiting. Seek medical advice immediately.

General Advice- No known delayed effects. Consult a physician for all exposures except for minor instance.

Section 5
Fire Fighting Measures
5.1 Flammability
The preparation is not flammable and is non-combustible.

5.2 Extinguishing Media
The product does not burn. Use a dry powder, foam or CO_2 type of fire extinguisher to fight the surrounding fire.

5.3 Combustion Products
When heated above 580°C, calcium hydroxide decomposes to produce calcium oxide (CaO) and water (H_2O). Calcium Oxide reacts with water and generates heat. This may cause risk to flammable material.

Section 6
Accidental Release Measures

6.1Personal Precautions
Avoid contact with skin and eyes

6.2Environmental Precautions
Contain the spillage. Avoid uncontrolled spills to watercourses and drains (pH rising). Any large spillage into watercourses must be alerted to the Environment Agency or other regulatory body.

6.3Methods for Cleaning Up
Pick up the product mechanically.

Section 7
Handling Ventilation and Storage

7.1Handling
Avoid contact with skin and eyes. Wear protective equipment (see section 8). When handling containers usual precautions should be paid to the risks outlines in the Council Directive 90/269EEC.

7.2Storage
Bulk storage should be in purpose designed tanks. Keep away from acids and nitro compounds. Keep out of reach of children. Do not use aluminium for transport or storage.

Section 8
Exposure Controls and Personal Protection

8.1Exposure Limit Values
CAS No/EINECS No – 1305-62-0/215-137-3 for the active substance $Ca(OH)_2$

Chemical Name - Suspension of Calcium Hydroxide in water

Occupational Exposure Standard (OES) – Not applicable

8.2Exposure Controls and Personal Protection
Occupational exposure controls – Not applicable

Respiratory protection – Not applicable

Hand Protection – Use approved nitrile impregnated gloves having CE marks

Eye protection – Tight fitting safety glasses with side shields or wide vision full goggles. Do not wear contact lenses when handling this product. It is also advisable to have individual pocket eyewash.

Skin protection – Use clothing fully covering skin, full length pants, long sleeved overalls with close fittings at openings. Footwear resistant to caustics.

General safety and hygiene measure – Wear clean, dry personal protective equipment.

Section 9
Physical and Chemicals Properties

Appearance:	White or off-white suspension in water
Odour:	Odourless
pH:	12.4 $Ca(OH)_2$ saturated solution at 25°C
Solubility:	1850 mg/1at 0°C for the active substance $Ca(OH)_2$
Melting Point:	For the active substance $Ca(OH)_2$, decomposition at 580°C to form CaO and H_2O
Boiling Point:	Not applicable
Specific gravity:	2.24 g/cm^3 at 20°C (active substance $Ca(OH)_2$)
Vapour pressure:	Non volatile

Section 10
Stability and Reactivity

10.1Condition to avoid
Minimise exposure to air to avoid degradation (carbonation). When heated above 580°C the active substance $Ca(OH)_2$ decomposes to produce CaO and H_2O.

10.2Materials to avoid
The active substance $Ca(OH)_2$ reacts with CO_2 $CaCO_3 + H_2O$
$Ca(OH)_2$ reacts with acids to form calcium salts.
Calcium Hydroxide reacts with aluminium and brass under formation (or release) of hydrogen gas:

$Ca(OH)_2 + 2Al + 6H_2O$ $Ca(Al(OH)_4)_2 + 3H_2$

10.3 Additional remarks
Calcium hydroxide absorbs carbon dioxide from air to form calcium carbonate, which is a common material in nature.

Section 11
Toxicological Information

11.1Acute effect:
Eye contact – risk of serious damage to eyes

Skin contact – Irritating to skin

Inhalation – Not applicable

Ingestion – The active substance, $Ca(OH)_2$, is not toxic but large amounts may cause irritation to the gastrointestinal tract.

11.2 Long Term exposure:
Eye contact – Risk of serious damage to eyes

Skin contact – **I**n case of prolonged skin contact, product may cause serious damage to skin

Inhalation – Not applicable

Section 12
Ecological Information

12.1Ecotoxicity
Acute/Prolonged toxicity to fish – On Gambusia affinis LC_{50} = 160 mg/l for 96 hours, the active substance, $Ca(OH)_2$ is non toxic because LC_{50} value is > 100 mg/l

Acute/Prolonged toxicity to aquatic invertebrates and plants – No test data

Toxicity to micro-organisms eg bacteria – At high concentration, through the rise of pH, calcium hydroxide is used for disinfection of sewage sludges.

Chronic toxicity to aquatic organisms – No data

Toxicity to soil dwelling organisms – No data

Toxicity to terrestrial plants – No data, however, calcium hydroxide is used as a fertiliser.

General effect – Acute pH-effect. Although the active substance $Ca(OH)_2$, is useful to correct water acidity, an excess of more than1g/l may be harmful to aquatic life. pH-value of >12 will rapidly decrease as a result of dilution and carbonation.

12.2Mobility
Calcium hydroxide reacts with carbon dioxide to form calcium carbonate, which is sparingly soluble, and so presents a low mobility in most ground. Moreover, this product is used as a fertiliser.

12.3Persistence and degradation
Not relevant for inorganic substances.

12.4Bioaccumulative Potential
Not relevant for inorganic substances.

Section 13
Disposal Consideration

Disposal should be in accordance with current local and national legislation.

Section 14
Transport Considerations

Classification:	Not classified as hazardous for transport
ADR (Road):	Not classified (Code UN3266 for Germany)
RID (Rail):	Not classified (Code UN3266 for Germany)
IMDG (Sea):	Not classified (Code UN3266 for Germany)
IATA (Air):	Code UN3266 – Class 8 – Packing Group III

Section 15
Regulatory Information

Symbol and classification of the substance according to Directive 67/548/EEC

Section 16
Other Information

Risk Phrases
R38: Irritating to skin
R41: Risk of serious damage to eyes

Safety Phrases
S2: Keep out of the reach of children
S25: Avoid contact with eyes
S26: In case of contact with eyes, rinse immediately with plenty of water and seek medical advice.
S37: Wear suitable gloves
S39: Wear eye/face protection

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